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$P_1V_1/T_1 = P_2V_2/T_2$ Chemistry Gas Laws Study Guide Flashcards | Quizlet

Ideal Gas Law. The last equation on that chart above is the ideal gas law, or $PV = nRT$. The

following information can be found on the AP Chemistry reference table but it's quick and easy to memorize! P = pressure in atm. V = volume in L. n = moles of gas. R = universal gas constant (0.08206

Latm/molK) T = temperature in Kelvin The Ideal Gas Law | Unit 3 - Intermolecular Forces and ... CHEMISTRY I (TESC 141)

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Avogadro's Law - the same number of moles (or molecules) of any gas

occupy the same volume at the same temperature and pressure Kinetic Energy of Gases $KE = \frac{1}{2} m v^2$ Where KE = kinetic

energy (in $kg \cdot m^2/s^2$ or "Joules") AP CHEMISTRY NOTES 5-1 THE GAS

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Boyle's Law states that the volume of a given mass of a dry gas is inversely proportional to its pressure at constant temperature. Charles' Law states that the volume of a given mass of a dry gas is directly proportional to its absolute (Kelvin) temperature if the pressure is kept constant.

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The pressure of a gas is directly proportional to the temperature if the volume remains constant What is the formula for the combined gas law?

$$P_1V_1/T_1 = P_2V_2/T_2$$

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